

m
(g)

V
(L)

$\frac{1\text{mole}}{\text{molar mass}}$

$\frac{\text{Molar mass}}{1\text{mole}}$

$\frac{22.4\text{L}}{1\text{mole}}$

$\frac{1\text{mole}}{22.4\text{L}}$

n
(mol)

$\frac{\text{Avagadro's}\#}{1\text{mole}}$

$\frac{1\text{mole}}{\text{Avagadro's}\#}$

of
Particles

Avogadro's # :

6.02×10^{23} particles / mol

Molar mass:

Calculated from Periodic Table
(g/mol)

MV:

@ STP = 22.4 L/mol